

Determination Of A Solubility Product Lab Answers

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get youtube without the ads. working no thanks 3 months free. find out why close. determination of the solubility product of an ionic compound lab explanation nathanjones0117. loading determination of a solubility product a laboratory report for chem 2145 performed by: kayla falcone performed: 3 october, 2013 date submitted: 17 october, 2013 abstract: the solubility of the salt calcium hydroxide, $\text{Ca}(\text{OH})_2$, is important to understand when determining the solubility product constant, K_{sp} determining a solubility product constant introduction the solubility product constant, K_{sp} , is the equilibrium constant for the solubility equilibrium of a slightly soluble (or nearly insoluble) ionic compound. calcium iodate is slightly soluble in distilled water at room temperature $\text{Ca}(\text{IO}_3)_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2 \text{IO}_3^{-}(\text{aq})$ best answer: which well was the first to show 0 precipitate? this is the well where $Q_{sp} < K_{sp}$ and the previous well is where $Q_{sp} > K_{sp}$ so, somewhere between these 2 wells is where $Q_{sp} = K_{sp}$, equilibrium well #7 is either at equilibrium, $Q_{sp} = K_{sp}$ or where $Q_{sp} < K_{sp}$. everything is still in solution. 7 well #7 1 determination of the solubility product of group II hydroxides introduction solubility equilibria many systems in chemistry appear to be static when in fact they are in (dynamic) equilibrium.

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